

Enhancement of Carbon Fiber Cloth Reactivity to Air by Potassium Addition using Ion Exchange

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Enhancement of the reactivity to air of polyacrylonitrile-based carbon fiber cloth has been achieved by the addition of metals to the cloth. Prior to metal addition, the cloth was oxidized in 54 wt% nitric acid in order, both, to increase the surface area of the cloth and to add carboxyl groups to the surface. Metal addition was then achieved by soaking the cloth in metal acetate solution to effect exchange between the metal cation and hydrogen on the carboxyl groups. The addition of potassium, sodium, calcium, and barium enhanced fiber cloth reactivity to air at 573 K. Extended studies using potassium addition showed that success in enhancing fiber cloth reactivity to air depends on: extent of cloth oxidation in nitric acid, time of exchange in potassium acetate solution, and the thoroughness of removing metal acetate from the fiber pore structure following exchange. With increased potassium loading on fiber cloth which has undergone a fixed extent of prior oxidation, cloth reactivity in air reaches a maximum and in some cases decreases. Optimum loading is associated with potassium added by exchange; additional loading is associated with potassium added by decomposition of acetate not removed from the fiber pore system by washing. Cloth reactivity increases essentially linearly with increase in potassium addition via exchange.

1. INTRODUCTION

Concern has arisen over the potential for accidental release of carbon fibers from resin matrix composites should the composite become involved in a fire¹). Once released from the resin matrix which is consumed in the fire, the light and fine carbon fibers can be carried in the atmosphere over large distances. Because of their high electrical conductivity, falling out of carbon fibers onto unprotected electrical and electronic circuits can cause a variety of problems, including resistive loading, temporary short circuits, and arcing. Thus for some applications there is interest in enhancing the reactivity of carbon fibers to air so that they will be consumed, along with the resin matrix, in a fire.

It is known that reactivity of carbon to oxidizing gases can be increased by increasing the active

surface area (ASA) of the carbon and by adding inorganic catalysts to the carbon^{2,3}). There have been a number of studies on increasing the surface area of carbon fibers using oxidative etching in order to subsequently enhance their bonding to the resin matrix⁴). Probably the use of nitric acid has been the most common⁴⁻⁷). Until recently, inorganics have been added to carbon primarily by impregnation of the carbon with a salt solution, followed by drying, and decomposition of the salt³). More recently, it has been shown that well dispersed inorganics (cations) can be effectively added to carbons using ion exchange with the hydrogen on carboxyl groups⁸). This has been the approach used in this study in an attempt to enhance fiber reactivity to air.

2. EXPERIMENTAL STUDIES

2.1 Materials

Studies were made on carbon fiber cloth woven by Woven Structures Inc. from T-300 carbon fiber yarn grade WYP 30 1/0 manufactured by Union

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Table 1 Selected properties of fibers in T-300 carbon cloth.

Surface Area, m ² /g	0.37
L_c , nm	1.1
Interlayer spacing, nm	0.356
Hg density, g/cc	1.71
X-ray density, g/cc	~2.14
He density, g/cc	1.79
Total Porosity, cc/g	0.117
Open Porosity, cc/g	0.026
Closed Porosity, cc/g	0.091

Carbide⁹). The polyacrylonitrile-based (PAN) yarn consisted of 3000 filaments (7 μ m in diameter) in a one-ply construction. The surface was treated with about 1.0–1.5 wt% proprietary sizing. Selected properties of the fibers in the cloth are given in Table 1. The small crystallite size (L_c) of the fiber cloth and its high nitrogen content (that is, 6.5 wt%) is consistent with its having been heated to less than 1800 K^{10,11}). Preferential basal plane alignment along the fiber axis results in a low surface area for the fibers, as measured from Kr adsorption at 77 K (using the BET equation) and a small amount of porosity accessible to helium at room temperature. In fact, the Kr surface area of 0.37 m²/g agrees well with the geometric area of the filaments (0.33 m²/g) based on a diameter of 7 μ m and density of 1.71 g/cc. It is significant that the fibers had a substantial total pore volume as calculated from x-ray and mercury densities, that is a total porosity of 20%, but that about 78% of the total pore volume was inaccessible to helium.

A semi-quantitative spectroscopic analysis of inorganic elements present in the fibers in the cloth is given in Table 2. Consistent with the findings of others¹²), the concentration of sodium* is found to be considerably higher than that of any other metal in the treated PAN fibers.

* In discussions with JPL personnel, they report finding wide variations in sodium content in PAN fibers. It is to be emphasized that our analysis was on just one batch of T-300 fiber cloth. Variations in sodium content in PAN may be due to such variables as: type of initiator used in polymerization, whether a sodium compound was present during the PAN spinning operation, and cation content in the wash water.

Table 2 Inorganic elements present in T-300 fiber cloth.

Element	Amount, wt%
Na	0.2
K	0.015
Ca	0.03
Ba	0.003
Fe	0.02
Mg	0.004
Zn	0.01
Ti	0.003
Al	0.01
Si	0.03

Not detected: B, Sb, Pb, Ge, In, Bi, V, Ni, Co, Be, Sr, Ce, La, Zr, Y, Mn, Cr, Ag, Cd, Sn.

2.2 Oxidative Etching

Different weights of cloth were refluxed in 54% HNO₃ solution (normality of 11.7) at about 388 K for 10 hr to: (i) increase the fiber accessible total surface area and ASA and (ii) add carboxyl groups to active carbon sites. Even for the highest cloth to solution ratio used (30 mg cloth/cc solution), the acid concentration of the solution was still high following oxidation; that is, it had a normality of 10.2. Oxidized samples were washed with distilled water under refluxing conditions for 40 hr and then dried at 383 K.

2.3 Ion Exchange

Replacement of hydrogen on the carboxyl groups of the HNO₃-oxidized fibers by metal cations was accomplished by shaking about 0.25 g of fiber cloth in a 50cc 2.0 molar solution of metal acetate in a N₂ atmosphere at 343 K for various periods of time. The procedure is described in more detail elsewhere¹³). Following exchange, the samples were washed either in flowing, cold distilled water or by refluxing in hot distilled water for varying periods of time.

2.4 Reactivity Measurements

Reactivity of fiber cloth samples to 0.1 MPa of dried air was measured using a Fisher Model 460 TGA unit. About 5 mg of sample were held in a platinum pan, suspended from a Cahn electrobalance. In all cases, samples were first heated to 383 K at a rate of 20 K/min in N₂ and held for 30 min to remove weakly adsorbed water. For

reactivity runs at 573 K, heating was continued in N_2 at 20 K/min up to temperature and held for 30 min before switching to air. Rates of burn-off at 573 K were expressed on a dry-ash free basis, that is following removal of water at 383 K. Gas flow rate through the reactor and sample weight in the platinum pan were such that an external gas mass transport effect on reactivity was absent²⁾.

3. RESULTS AND DISCUSSION

3.1 Nitric Acid Oxidized Fiber Cloth

Oxidation of carbon fiber cloth in nitric acid decreased the level of inorganic impurities, increased surface area, reduced cloth reactivity to air, but produced a negligible change in L_c or interlayer spacing within the crystallite. For example, oxidation for 10 hr, using a cloth to HNO_3 ratio of 6.0 mg/cc, reduced the sodium concentration from 0.2 to 0.03 wt%, calcium below 0.03%, and barium below 0.003%.

As seen in Table 3, surface area development depends upon the ratio of weight of fiber cloth oxidized to volume of acid used for a 10 hr oxidation time. Surface areas were measured using Kr adsorption at 77 K and the BET equation with 30 min allowed for equilibration of each adsorption point. Outgassing temperatures between 383 and 573 K prior to adsorption had a negligible effect on Kr adsorption. The particularly large increase in surface area in going from a cloth to acid ratio of 6.6 to 6.0 mg/cc was accompanied by a significant change in the chemical analysis of the fiber cloth. That is, the original cloth and those oxidized using cloth to acid ratios of 6.6 and greater had carbon contents between 93.2 and 88.4%, nitrogen contents between 6.3 and 6.8%, negligible hydrogen and sulfur, and the remainder oxygen and water. The cloth oxidized using a cloth to acid ratio of 6.0 had the following analysis: C, 70.9%; N, 4.9%; H, 0.73%; and balance

Table 3 Surface area development in T-300 fiber cloth upon HNO_3 oxidation.

Sample	Cloth to Acid mg/cc	Surface Area m^2/g
A	30	1.0
B	25	1.2
C	6.6	1.5
D	6.0	2.3

23.5%. Obviously there is a large increase in oxygen and water content in this sample; but oxidation added negligible additional nitrogen to the cloth, for example as nitro groups.* Heating the sample to 383 K in N_2 resulted in a weight loss of 11% which is attributed primarily to release of water which was hydrogen bonded to oxygen functional groups on the fiber cloth surface. Further heating of the sample to 573 K in N_2 resulted in an additional weight loss of 10% of the dried fiber cloth, which is attributed primarily to loss of CO_2 following decomposition of carboxyl groups from the fiber surface¹⁵⁾.

Assuming that 12.5% of the oxidized fiber cloth (Sample D) is oxygen in the form of carboxyl groups and that each carboxyl group occupies an area of 0.083 nm^2 (associated with an active site in the prismatic surface of a carbon crystallite), the area occupied by carboxyl groups is estimated to be $195 \text{ m}^2/g$. This area is considerably in excess of that measured by Kr adsorption at 77 K on this sample and suggests molecular sieving¹⁶⁾. That is, HNO_3 at 388 K was able to penetrate significant pore volume which was inaccessible to Kr at 77 K. Or possibly HNO_3 gained access by swelling the fiber. That oxidized fiber cloth sample D was a molecular sieve was further confirmed by measuring CO_2 uptake at 298 K. From the Dubinin-Polanyi plot¹⁷⁾, a micropore surface area of $139 \text{ m}^2/g$ is estimated, again a value considerably in excess of the area measured by Kr adsorption at 77 K.

Reactivity of the as-received fiber cloth in air at 573 K was very low; that is less than 1% weight loss in 10 hr and less than 4% weight loss in 70 hr. Reactivity of the HNO_3 -oxidized sample D was still lower, that is 0.6% in 10 hr. Presumably this decrease in reactivity can be attributed to reduction in inorganic impurity content over that of the as-received cloth.

3.2 Ion Exchanged Fiber Cloth

In exploratory studies, potassium, sodium, calcium, and barium were exchanged onto the

* The latter finding is consistent with results of Deno who reports that a significant number of nitro groups are added to aromatic systems only when reacted with nitric acid at concentrations considerably in excess of the 54 wt% used in this study¹⁴⁾.

e increase in sample; but all nitrogen to 100%.* Heating in a weight loss experiment to release oxygen from the cloth surface. 573 K in N_2 for 10% of the weight loss is due primarily to the loss of carboxyl

zed fiber cloth. The amount of carboxyl groups occupies an active site on crystallite), the weight loss is estimated to be in excess of 10% at 77 K on cooling¹⁶). That rate is significant compared to Kr at 77 K. The weight loss by swelling the sample D was also measured by measuring the Dubinin surface area of 139 m²/g, considerably in excess of adsorption at

er cloth in air at 573 K. The weight loss is more than 1% weight loss in 70 hr. The weight loss of sample D was also measured, presumably this is due to the reduction of potassium over that of

assium, sodium, and magnesium onto the

with results of the present experiment. A significant number of active systems only at low concentrations of potassium (0.04 wt% used in

fiber cloth surface. As expected^{3,18}), they all catalyzed cloth reactivity to air at 573 K. However, for each cation, reactivity did not increase monotonically with increased loading; instead it went through a maximum. This would not be expected if the metal was added only by ion exchange⁸). It was thought that some metal may have been derived from the metal acetate solution which was not completely removed from the cloth pore structure by washing. A more thorough study was performed using potassium, which is described herein, in an attempt to understand relations between cloth reactivity and how the metal was added.

In the first set of experiments, the cloth which had been oxidized using a cloth to HNO_3 ratio of 30 mg/cc (Sample A) was exchanged in a potassium acetate solution (starting pH of 8.4) for 50 hr. The sample was then washed by refluxing in hot distilled water for varying lengths of time to remove potassium from the pore system. Figure 1 presents results for the amount of potassium remaining on the cloth as a function of washing time. It is clear that metal loading will be a function of not only the extents of oxidation in HNO_3 and subsequent cation exchange but also the severity of final washing. The curve in Figure 1 is characteristic of two concurrent processes of significantly different rates. It is suggested that the rapid process involves removal of the acetate

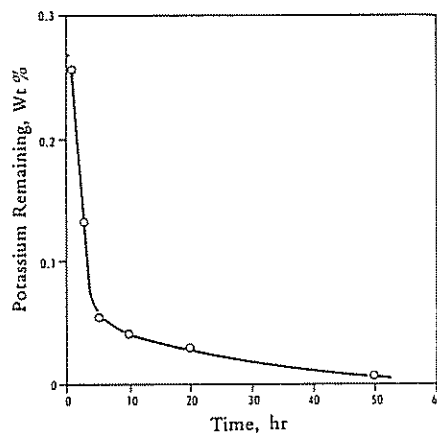


Figure 1. Effect of time of washing by refluxing in hot water on potassium content in cloth sample A. Sample A previously exchanged with potassium acetate for 50 hr.

from the pore system and the slow process involves back-ion exchange of hydrogen from hot water replacing potassium on the carboxyl groups. Extrapolation of the lower curve back to the ordinate yields a potassium concentration on the cloth at zero time of about 0.07%, which is taken as the maximum value of cation loading under the conditions of previous cloth oxidation and ion exchange used. It is concluded that following washing by refluxing for about 5 hr essentially all potassium derived from the acetate has been removed from the pore system and the potassium remaining is associated with carboxyl groups. For refluxing times progressively shorter than 5 hr, the contribution of potassium derived from the acetate to total potassium present increases sharply.

If we reason that the rate of back-ion exchange is more temperature dependent (that is, has a higher activation energy for its rate constant) than the rate of washing of the acetate from the pore system, then the use of cold water should separate the effects of these two processes still better than hot water. Thus washing of the cation exchanged cloth was also conducted in a continuous flow of fresh cold water (where an acetate ion concentration was not allowed to build up in the wash water as with refluxing). Following washing in cold water for 1 hr, the potassium concentration on the cloth was 0.078%, suggesting that essentially all acetate was washed out of the pore system but that little potassium cation associated with carboxyl groups had been removed.

Reactivity plots for the fiber cloth samples containing amounts of potassium varying from 0.006 to 0.265% are shown in Figure 2 for reaction times up to 8 hr. In some cases duplicate runs were made; in these cases data points give the average values and bars give the spread in values. Even though there is a monotonic increase in reactivity with increasing potassium loading, it is obvious that the specific activity of the catalyst decreases at higher loadings, as seen in Figure 3 where burn-off following reaction for 8 hr is plotted against potassium loading. This could be due to at least three reasons: (i) decreasing dispersion of the exchanged potassium with increasing concentration, (ii) poorer dispersion of the potassium derived from potassium acetate than that derived from the exchanged cation, and (iii) blockage of

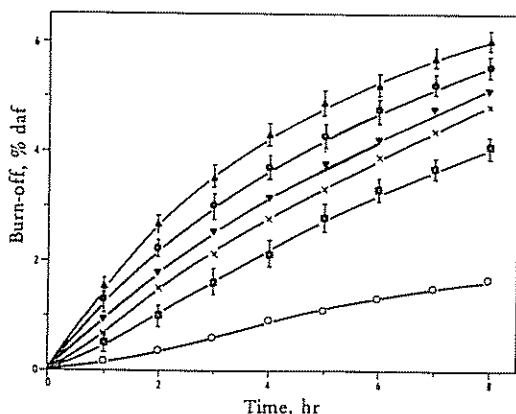


Figure 2. Reactivity plots of cloth sample A in air at 573 K for various potassium loadings, in wt%: ○, 0.006; ■, 0.031; ×, 0.045; ▼, 0.051; ●, 0.133; ▲, 0.265.

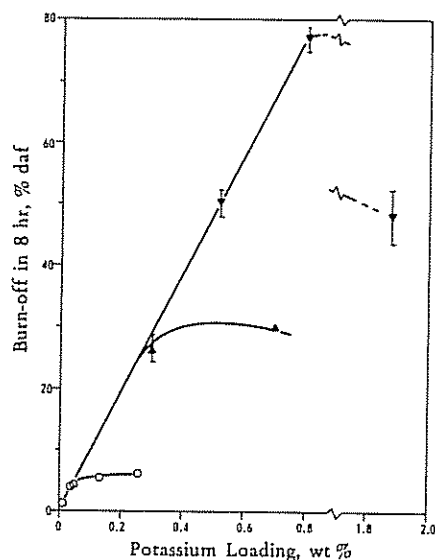


Figure 3. Extent of burn-off in 8 hr in air at 573 K of cloth samples loaded with varying amounts of potassium. Sample: ○, A; ▲, C; ▼, D.

some of the pore system. Krypton surface areas of samples loaded to 0.133 and 0.265% potassium actually increased ($1.9 \text{ m}^2/\text{g}$) over that of the oxidized fiber ($1.0 \text{ m}^2/\text{g}$), and thus pore blockage at this level of loading was not present. The possibility that the degree of dispersion of the metal (or its oxide) is different following decomposition of the metal carboxylate which is bonded to the surface compared to decomposition of

metal acetate particles (agglomerates) sitting on the surface is interesting to contemplate. A definitive answer does not come from this study, but is an important question when considering catalyst addition to carbon substrates and subsequent catalyst efficiency.

In additional experiments, cloths oxidized using cloth to HNO_3 ratios of 6.6 and 6.0 mg/cc were exchanged with potassium acetate. For sample C, exchange was conducted for 30 hr, followed by washing for 1 hr using either refluxing in hot water or flowing cold water. Concentrations of potassium remaining in the fiber were 0.71 and 0.31%, respectively. It is suggested, as just discussed, that the loading of 0.30% essentially represents potassium derived from carboxyl groups and that the loading of 0.71% represents potassium derived both from carboxyl and acetate entities. Enhanced loading from these two sources onto cloth sample C over that found on exchanged cloth sample A is consistent with the larger surface area (and pore volume) in the former sample.

The reactivity plot for the 0.30% potassium loaded sample is shown in Figure 4. Over the 8 hr reaction period, reactivity continuously decreases with increasing reaction time despite a build-up in catalyst concentration in the carbon remaining. This suggests that: (i) increasing agglomeration (sintering) of catalyst particles is occurring with increasing burn-off, that is catalyst dispersion is decreasing and/or (ii) the catalyst is losing contact with the carbon surface. McKee and Chatterji have shown that the mechanism of catalysis of carbon gasification by potassium involves oxidation-reduction steps¹⁹. That is, the higher oxide K_2O_2 is reduced by the carbon to K_2O which, in turn, is reoxidized to K_2O_2 by air. The rate of agglomeration is dependent upon the mobility of the catalyst species on the surface. Even though reaction temperature was below the melting point of K_2O_2 (663 K), Thomas and Walker²⁰ have shown that solid catalyst particles can have high mobility on carbon surfaces, particularly in an oxygen environment. Further, Baker has shown, from in-situ electron microscopy studies, that species on carbon surfaces in an O_2 atmosphere commence having significant mobility above their Tamman temperature, that is one half their melting temperature, expressed in degrees K²¹.

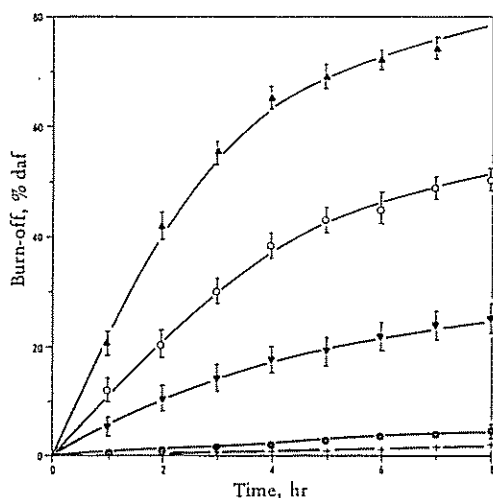


Figure 4. Reactivity plots of cloth samples with various potassium loadings in air at 573 K.

Sample: +, A-0.006% K; ⊙, A-0.031% K; ▼, C-0.30% K; ○, D-0.52% K; ▲, D-0.79% K.

Thus, at a reaction temperature of 573 K, K_2O_2 (having a Tammann temperature of 332 K) would be expected to have significant mobility on carbon surfaces.

Burn-offs in 8 hr for the samples loaded to 0.30 and 0.71% potassium are given in Figure 3. As with cloth sample A, it is suggested that there is an optimum potassium loading to achieve maximum cloth reactivity for Sample C. Additional loading increases reactivity only slightly or may, in some cases, even decrease reactivity, as will now be seen.

For sample D, exchange was conducted for 10, 30, and 50 hr in potassium acetate followed by washing in running cold water for 1 hr. Loadings of potassium were 0.52, 0.79, and 1.8%, respectively. These results suggest that transport of the acetate solution into and out of the fiber cloth pore system was slow. This is consistent with this oxidized sample having some molecular sieve character, as previously discussed. Figure 4 presents reactivity plots for the samples loaded to 0.52 and 0.79% potassium. Again reactivities decrease continually with increasing time of reaction. Burn-offs in 8 hr for all samples are given in Figure 3. For sample D, again an optimum potassium loading is suggested; at higher loadings there is a significant

decrease in cloth reactivity, probably due to pore blockage and/or coverage of the carbon surface with more than a monolayer of potassium. That is, washing the sample, which was exchanged for 50 hr, in running cold water for 1 hr appears to have not removed all potassium acetate from the pore system. If optimum loading is taken as about 0.79% potassium and if it is assumed that each potassium atom was previously associated with a carboxyl group, it is calculated that the potassium carboxylate occupied about $10 \text{ m}^2/\text{g}$ of surface. Again this is larger than the Kr area of oxidized cloth sample D and is consistent with molecular sieving. However, the area is much less than that which was estimated to be occupied by carboxyl groups following nitric acid oxidation (that is, $195 \text{ m}^2/\text{g}$). This suggests that: (i) potassium only exchanged on a small fraction of the carboxyl groups following oxidation, (ii) oxidation in nitric acid added both carboxyl, acidic phenolic and lactone groups to the fiber surface, and/or (iii) nitric acid was not completely washed from the pore system.

4. CONCLUSIONS

The reactivity of PAN fiber cloth to air can be substantially increased by the addition of metals to the cloth via ion exchange from metal acetate solutions. However, since the surface area of PAN fibers (even those heated to less than 1800 K) is very low, significant metal addition via ion exchange can only be achieved if the fiber is first treated with nitric acid to increase surface area and, as importantly, to add carboxyl groups. The amount of metal which can be added to the fiber is a function of the severity of oxidative treatment, the time of subsequent exchange, and the degree of subsequent removal of metal from the fiber by washing. Fiber cloth reactivity to air appears to increase essentially linearly with increase in potassium loading via ion exchange. Additional potassium loading originating from potassium acetate left in the pore system of the fibers following washing appears to increase reactivity only slightly in some cases and in other cases to, in fact, result in reduction in fiber cloth reactivity. It is suspected that there will be an optimum in the extent of fiber oxidation to achieve enhanced reactivity. At the optimum, both fiber cloth

reactivity and mechanical properties of composites made therefrom will be enhanced. Too much fiber oxidation will lead to an unacceptable reduction in mechanical properties of the composites subsequently fabricated, even though fiber reactivity to air may be further enhanced.

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REFERENCES

- 1) S. E. Wentworth, A. O. King, and R. J. Shuford, Army Materials & Mechanics Research Center Report No. TR79-1, January, 1979.
- 2) P. L. Walker, Jr., F. Rusinko, Jr., and L. G. Austin, *Advances in Catalysis* (Edited by D.D. Eley, P.W. Selwood, and P. B. Weisz), Vol. 11, pp. 133–221, Academic Press, New York (1959).
- 3) P. L. Walker, Jr., M. Shelef, and R. A. Anderson, *Chemistry and Physics of Carbon* (Edited by P. L. Walker, Jr.), Vol. 4, pp. 287–380, Marcel Dekker, New York (1968).
- 4) D. W. McKee and V. J. Mimeault, *Chemistry and Physics of Carbon* (Edited by P. L. Walker, Jr. and P. A. Thrower), Vol. 8, pp. 151–245, Marcel Dekker, New York (1978).
- 5) C. S. Brooks, G. S. Golden, and D. A. Scola, *Carbon* 12, 609 (1974).
- 6) B. Rand and R. Robinson, *Carbon* 15, 257 (1977).
- 7) B. Rand and R. Robinson, *Carbon* 15, 311 (1977).
- 8) E. J. Hippo, R. G. Jenkins, and P. L. Walker, Jr., *Fuel* 58, 338 (1979).
- 9) Technical Information Bulletin No. 465–223, Union Carbide Corp.
- 10) S. C. Bennett and D. J. Johnson, *Carbon* 17, 25 (1979).
- 11) W. Watt, *Carbon* 10, 121 (1972).
- 12) L. T. Drzal, *Carbon* 15, 129 (1977).
- 13) H. N. S. Schafer, *Fuel* 49, 197 (1970).
- 14) N. C. Deno, The Pennsylvania State University, Private Communication, November, 1979.
- 15) H. N. S. Schafer, *Fuel* 58, 673 (1979).
- 16) P. L. Walker, Jr., L.G. Austin, and S. P. Nandi, *Chemistry and Physics of Carbon* (Edited by P. L. Walker, Jr.), Vol. 2, pp. 257–371, Marcel Dekker, New York (1966).
- 17) M. M. Dubinin, *Chemistry and Physics of Carbon* (Edited by P. L. Walker, Jr.), Vol. 2, pp. 51–120, Marcel Dekker, New York (1966).
- 18) P. L. Walker, Jr., O. P. Mahajan, and M. Komatsu, Preprints American Chemical Society, Division Fuel Chemistry Meeting, Vol. 24, No. 3, 10 (1979).
- 19) D. W. McKee and D. Chatterji, *Carbon* 13, 381 (1975).
- 20) J. M. Thomas and P. L. Walker, Jr., *J. Chem. Phys.* 41, 587 (1964).
- 21) R. T. K. Baker, Exxon Corp., Private Communication, January, 1980.